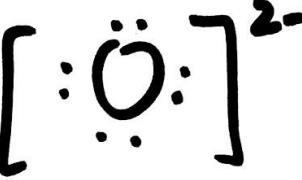
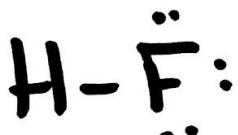
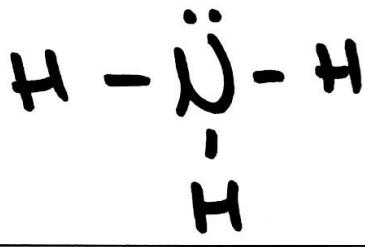
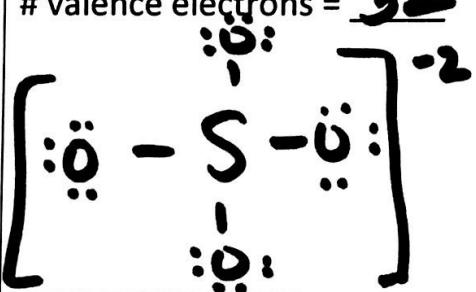
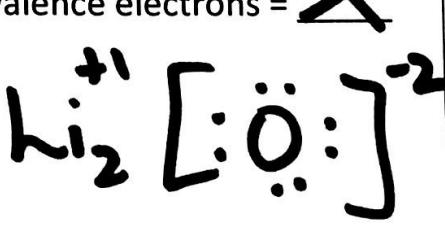
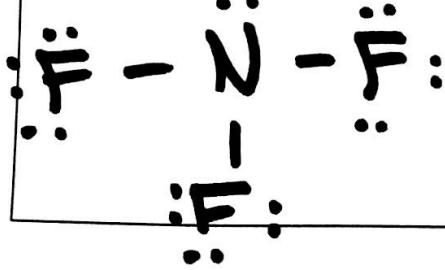
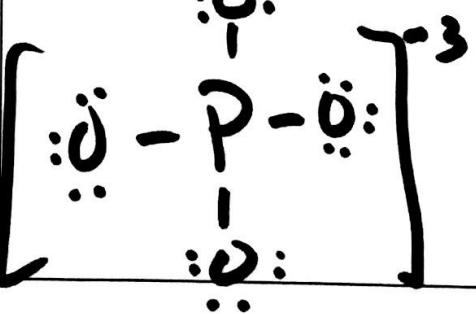


Chemical Bonding

Indicate the number of valence electrons for each. Write the correct Lewis electron dot structure for each.

Lewis Dot Structures

F # valence electrons = <u>7</u> 	O # valence electrons = <u>6</u> 	K # valence electrons = <u>1</u> 
F ⁻¹ # valence electrons = <u>8</u> 	O ²⁻ # valence electrons = <u>8</u> 	Al ³⁺ # valence electrons = <u>0</u> 
HF # valence electrons = <u>8</u> 	CaCl ₂ # valence electrons = <u>X</u> 	NH ₃ # valence electrons = <u>8</u> 
SO ₄ ²⁻ # valence electrons = <u>32</u> 	CO # valence electrons = <u>10</u> 	Li ₂ O # valence electrons = <u>X</u> 
NF ₃ # valence electrons = <u>26</u> 	PO ₄ ³⁻ # valence electrons = <u>32</u> 	CBr ₄ # valence electrons = <u>32</u> 