

Chemical Bonding

Lewis Dot Structures

Indicate the number of valence electrons for each. Write the correct Lewis electron dot structure for each.

<p>F # valence electrons = <u>7</u></p> <p>$\cdot\ddot{\text{F}}\cdot$</p>	<p>O # valence electrons = <u>6</u></p> <p>$\cdot\ddot{\text{O}}\cdot$</p>	<p>K # valence electrons = <u>1</u></p> <p>K\cdot</p>
<p>F⁻¹ # valence electrons = <u>8</u></p> <p>$[\cdot\ddot{\text{F}}\cdot]^{-1}$</p>	<p>O²⁻ # valence electrons = <u>8</u></p> <p>$[\cdot\ddot{\text{O}}\cdot]^{2-}$</p>	<p>Al³⁺ # valence electrons = <u>0</u></p> <p>Al</p>
<p>HF # valence electrons = <u>8</u></p> <p>H-$\ddot{\text{F}}\cdot$</p>	<p>CaCl₂ # valence electrons = X</p> <p>Ca⁺² [$\ddot{\text{Cl}}\cdot$]₂⁻¹</p>	<p>NH₃ # valence electrons = <u>8</u></p> <p>H-$\ddot{\text{N}}\cdot$-H H</p>
<p>SO₄²⁻ # valence electrons = <u>32</u></p> <p>$[\cdot\ddot{\text{O}}-\text{S}(\ddot{\text{O}})_3\cdot]^{2-}$</p>	<p>CO # valence electrons = <u>10</u></p> <p>$:\text{C}\equiv\text{O}:$</p>	<p>Li₂O # valence electrons = X</p> <p>Li⁺¹ [$\ddot{\text{O}}\cdot$]₂⁻²</p>
<p>NF₃ # valence electrons = <u>26</u></p> <p>$\cdot\ddot{\text{F}}-\ddot{\text{N}}(\ddot{\text{F}})_2\cdot$</p>	<p>PO₄³⁻ # valence electrons = <u>32</u></p> <p>$[\cdot\ddot{\text{O}}-\text{P}(\ddot{\text{O}})_3\cdot]^{3-}$</p>	<p>CBr₄ # valence electrons = <u>32</u></p> <p>$\cdot\ddot{\text{Br}}-\text{C}(\ddot{\text{Br}})_3\cdot$</p>