|  |  |
| --- | --- |
| Question:  **-245oC converted to Kelvin** | Question:  How many Joules are absorbed by 5.0 g of water is heated from 10.0oC to 25.0oC |
| Question:  How much energy is required to vaporize 15.0g of water at 100oC | Question:  How much energy is required to completely melt 50.0 g of water at its melting point? |
| Question:  Convert 373 K to oC | Question:  A 20.0 g sample of H2O(l) at 25.0oC absorbs 334 joules of heat. What is the final temperature of the H2O(l) sample? |
| Question:  The temperature of a sample of water changes from 10.0oC to 20oC when the sample absorbs 878 joules of heat. What is the mass of the sample? | Question:  What is the maximum number of kilojoules needed to change 25.0 g of water at 100.0oC to steam? |
| Question:  What is the total amount of heat energy released when 50.0 grams of water cools from 50.0°C to 35.0°C. | Question:  What is the maximum number of kilojoules needed to change 45.0 g of H2O(s) to H2O(l) at its freezing point? |
| Answer:  28 | Answer:  15 |
| Answer:  33,900 | Answer:  310 |
| Answer:  100 | Answer:  16,700 |
| Answer:  21 | Answer:  29 |
| Answer:  3,140 | Answer:  56.5 |