|  |  |
| --- | --- |
| Question: **-245oC converted to Kelvin** | Question:How many Joules are absorbed by 5.0 g of water is heated from 10.0oC to 25.0oC |
| Question:How much energy is required to vaporize 15.0g of water at 100oC | Question:How much energy is required to completely melt 50.0 g of water at its melting point? |
| Question:Convert 373 K to oC | Question:A 20.0 g sample of H2O(l) at 25.0oC absorbs 334 joules of heat. What is the final temperature of the H2O(l) sample? |
| Question:The temperature of a sample of water changes from 10.0oC to 20oC when the sample absorbs 878 joules of heat. What is the mass of the sample? | Question:What is the maximum number of kilojoules needed to change 25.0 g of water at 100.0oC to steam? |
| Question:What is the total amount of heat energy released when 50.0 grams of water cools from 50.0°C to 35.0°C. | Question:What is the maximum number of kilojoules needed to change 45.0 g of H2O(s) to H2O(l) at its freezing point? |
| Answer:28 | Answer:15 |
| Answer:33,900  | Answer:310  |
| Answer:100 | Answer:16,700  |
| Answer:21 | Answer:29 |
| Answer:3,140 | Answer:56.5 |